

班级\_\_\_\_\_姓名\_\_\_\_\_号数\_\_\_\_\_成绩\_\_\_\_\_

1. “ ”  
 A. B.  
 C. D. “ ”
2. A. B. C. D.
3. A.H<sub>2</sub>O B.Na<sub>2</sub>O C.NaOH D. NaCl
4. A. B. C. D.
5. A. 1.5 mol H<sub>2</sub> B. 44.8LSO<sub>2</sub>  
 C.96gO<sub>2</sub>(O<sub>2</sub> 32g·mol<sup>-1</sup>) D.N<sub>A</sub> N<sub>2</sub>(.N<sub>A</sub> )
6. A.H<sup>+</sup> Na<sup>+</sup> OH<sup>-</sup> B.Fe<sup>3+</sup> OH<sup>-</sup> Cl<sup>-</sup> C.Cu<sup>2+</sup> SO<sub>4</sub><sup>2-</sup> NO<sub>3</sub> D.H<sup>+</sup> Ca<sup>2+</sup> HCO<sub>3</sub><sup>-</sup>
7. :CH<sub>2</sub>=CH<sub>2</sub>+Br<sub>2</sub>⇌CH<sub>2</sub>BrCH<sub>2</sub>Br  
 A. B. C. D.
8. A.O<sub>2</sub> O<sub>3</sub> B.<sup>1</sup>H <sup>2</sup>H C.CO CO<sub>2</sub> D.CH<sub>3</sub>CH<sub>3</sub> CH<sub>3</sub>CH<sub>2</sub>CH<sub>3</sub>
- 9.“ ” “ ” “ ” “ ”  
 A. B. C. D.
10. 250mL0.1ml·L<sup>-1</sup> Na<sub>2</sub>CO<sub>3</sub>



A.



B.



C.



D.


11. A.NH<sub>4</sub>Cl⇌NH<sub>3</sub>↑+HCl↑ B.SO<sub>2</sub>+ 2NaOH⇌ Na<sub>2</sub>SO<sub>3</sub>+H<sub>2</sub>O  
 C.SO<sub>3</sub>+H<sub>2</sub>O⇌H<sub>2</sub>SO<sub>3</sub> D.2Na +2H<sub>2</sub>O⇌2NaOH+ H<sub>2</sub>↑
12. ,  
 A. B.  
 C. D.
13. K<sub>2</sub>SO<sub>4</sub> NH<sub>4</sub>Cl (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> ,  
 A.Ba(OH)<sub>2</sub> B.BaCl<sub>2</sub> C.NaOH D.AgNO<sub>3</sub>
14. A. B.  
 C. D.

15.

- (1) \_\_\_\_\_
- (2) \_\_\_\_\_ (3) O S \_\_\_\_\_
- (4) \_\_\_\_\_ (5) \_\_\_\_\_
- (6) Na Mg \_\_\_\_\_
- (7) NH<sub>3</sub> PH<sub>3</sub> \_\_\_\_\_
- (8) H<sub>2</sub>SO<sub>4</sub> H<sub>3</sub>PO<sub>4</sub> \_\_\_\_\_
- (9) Mg(OH)<sub>2</sub> Al(OH)<sub>3</sub> , \_\_\_\_\_
- (10) \_\_\_\_\_ )

22.(6 ) ,

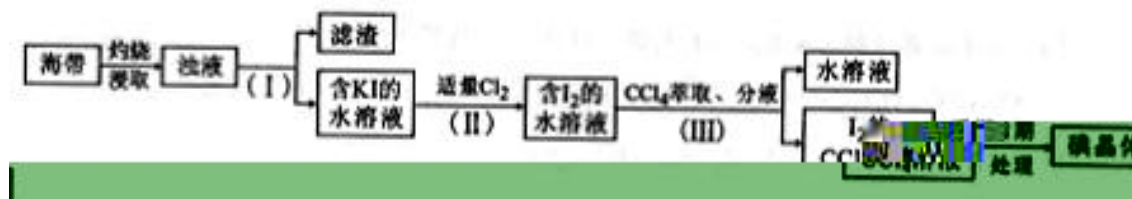
A. B. C. D. E. F.

- (1) \_\_\_\_\_ (2) \_\_\_\_\_
- (3)  \_\_\_\_\_ (4) \_\_\_\_\_
- 5) \_\_\_\_\_ (6) \_\_\_\_\_

23.(6 )

- (1) \_\_\_\_\_
- (2) \_\_\_\_\_
- (3) \_\_\_\_\_

24.(6 ) , :

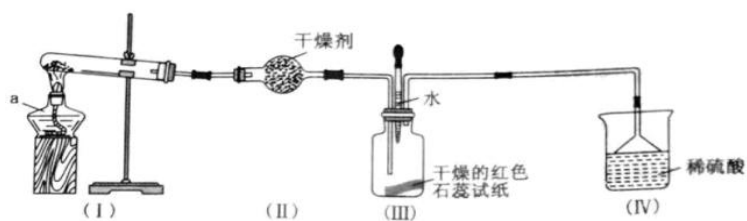
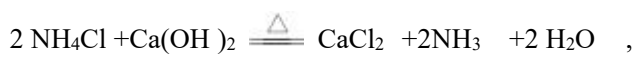


- (1) (I) \_\_\_\_\_
- (2) (II) KI Cl<sub>2</sub> \_\_\_\_\_ , \_\_\_\_\_
- (3) (III) I<sub>2</sub> CCl<sub>4</sub> \_\_\_\_\_ ( “ ” )

25.(15 )

NH<sub>3</sub>

[



(1) (I) a \_\_\_\_\_  
 (2) (II) \_\_\_\_\_  
 (3) \_\_\_\_\_ (III) \_\_\_\_\_ ) \_\_\_\_\_ )  
 (4) \_\_\_\_\_ III \_\_\_\_\_ ) , NH<sub>3</sub> HCl NH<sub>4</sub>Cl

(5) (IV) \_\_\_\_\_  
 (6) 0.1mol NH<sub>4</sub>Cl Ca(OH)<sub>2</sub> NH<sub>3</sub>  
 \_\_\_\_\_ mol

(7) 2 NH<sub>3</sub> + 3CuO  $\xrightarrow{\Delta}$  3Cu + N<sub>2</sub> + 3 H<sub>2</sub>O, III  
 \_\_\_\_\_ ) NH<sub>3</sub> \_\_\_\_\_  
 \_\_\_\_\_ )

(8) \_\_\_\_\_

26. (15

① 已知C(石墨, s) + O<sub>2</sub>(g) = CO<sub>2</sub>(g) ΔH = -393.5 kJ·mol<sup>-1</sup>  
 H<sub>2</sub>(g) +  $\frac{1}{2}$  O<sub>2</sub>(g) = H<sub>2</sub>O(g) ΔH = -241.8 kJ·mol<sup>-1</sup>

(2) CO<sub>2</sub>  
 CO<sub>2</sub>(g) + 3H<sub>2</sub>(g) = CH<sub>3</sub>OH(g) + H<sub>2</sub>O(g) H = -49.0 kJ·mol<sup>-1</sup>  
 CO<sub>2</sub> H<sub>2</sub>, CO<sub>2</sub>(g)  
 CH<sub>3</sub>OH(g)  
 v(H<sub>2</sub>) = \_\_\_\_\_  
 \_\_\_\_\_ mol·L<sup>-1</sup>·min<sup>-1</sup>

A. \_\_\_\_\_ ) B. CO<sub>2</sub> C.

(3) CH<sub>3</sub>COOH