

1. ()

- A
- B
- C
- D

2. 1 mol 241.8 kJ 1 g 2.45 kJ ()

- A $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l}) \quad \Delta \quad 285.9 \text{ kJ}\cdot\text{mol}^{-1}$
- B $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l}) \quad \Delta \quad 241.8 \text{ kJ}\cdot\text{mol}^{-1}$
- C $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l}) \quad \Delta \quad 285.9 \text{ kJ}\cdot\text{mol}^{-1}$
- D $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g}) \quad \Delta \quad 241.8 \text{ kJ}\cdot\text{mol}^{-1}$

3. ()

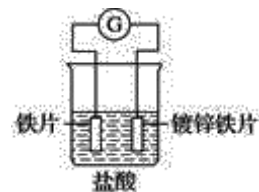
- A HNO_3 pH
- B KOH OH pH
- C K_2SO_4 1 2
- D CuCl_2 1 1

4. I II ()

选项	陈述 I	陈述 II	判断
A	用锌和足量稀硫酸制取氢气时加入硫酸铜溶液	可以制取更多的氢气	I对、II对、有
B	用 Mg-Al-NaOH 构成原电池	Mg 更活泼作负极	I对、II对、无
C	AlCl_3 是离子化合物	电解熔融的 AlCl_3 制取金属铝	I对、II对、无

5. ()

- A pH
- B
- C $\text{Fe} - 2\text{e} \rightleftharpoons \text{Fe}^{2+}$
- D



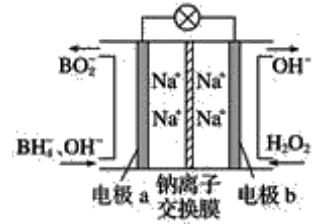
6. $\text{NaBH}_4(\text{B} \quad 3) \quad \text{H}_2\text{O}_2$
 Pt/C MnO_2 ()

A Na b a

B b MnO₂ MnO₂

C 1 mol H₂O₂ 1 mol

D $\text{BH}_4^- + 8\text{OH}^- - 8\text{e}^- \rightleftharpoons \text{BO}_2^- + 6\text{H}_2\text{O}$



7. — $2\text{Li(s)} + \text{I}_2(\text{s}) \rightleftharpoons 2\text{LiI(s)} \quad \Delta$

$4\text{Li(s)} + \text{O}_2(\text{g}) \rightleftharpoons 2\text{Li}_2\text{O(s)} \quad \Delta_1$ $4\text{LiI(s)} + \text{O}_2(\text{g}) \rightleftharpoons 2\text{I}_2(\text{s}) + 2\text{Li}_2\text{O(s)} \quad \Delta_2$

()

A $\Delta - \frac{1}{2}\Delta_1 - \Delta_2$ B $\Delta - \frac{1}{2}\Delta_1 - \Delta_2$ C $\Delta - \frac{1}{2}\Delta_1 - \frac{1}{2}\Delta_2$ D $\Delta - \frac{1}{2}\Delta_1 - \frac{1}{2}\Delta_2$

8. ()

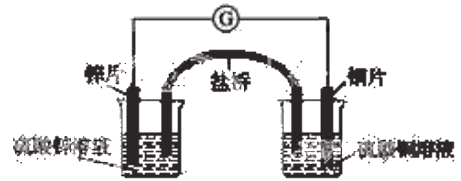
A

B

C

D $2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2\uparrow$

D



9. A B NaCl

B c

d ()

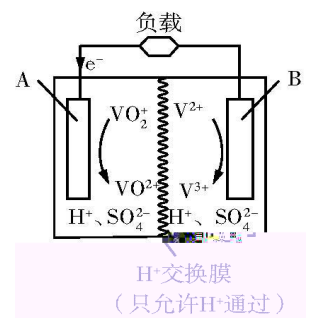
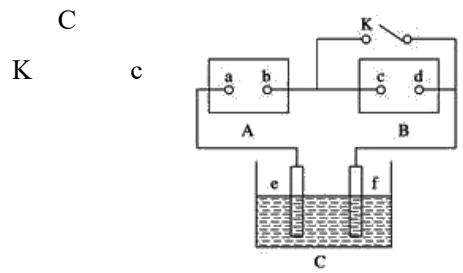
A a

B c $2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2\uparrow$

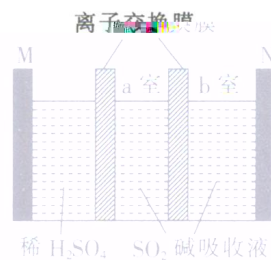
C f

D e

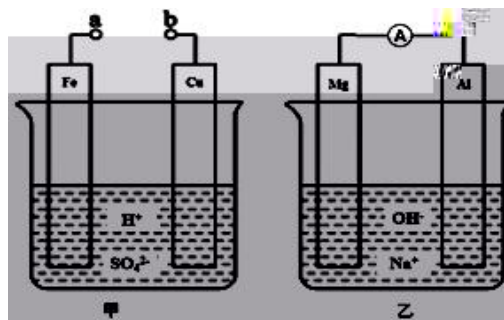
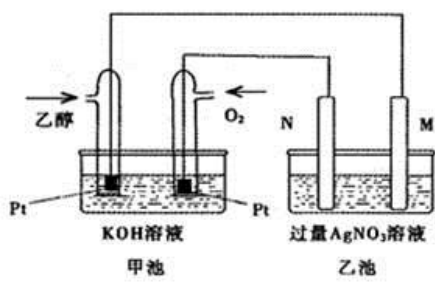
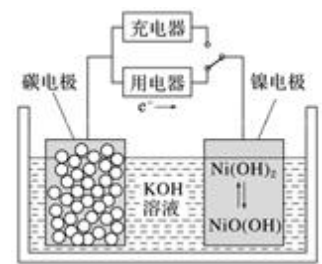
10.



11.



12.



1 () O_2

()

2 NO_3^- (N) 4.32g

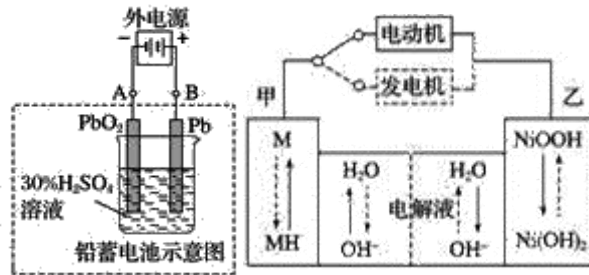
L()





2H₂O

()



(1)

H₂SO₄

1 mol

g

(2)

PbO₂ Pb

A

B

16.

(1)

(C₈H₁₈)

1 mol

569.1 kJ

(2)

(M)

(KOH)

H₂



pH (“ ” “ ” “ ”)

(3) Cu₂O

2Cu



1 mol H₂

Cu₂O

g



N 0.224

